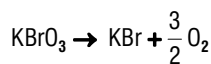


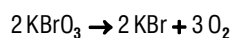
GABARITO COMENTADO

QUÍMICA

01. Letra B.



ou

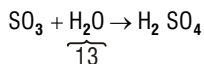
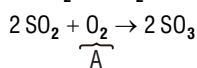
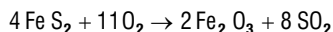


02. Letra E.

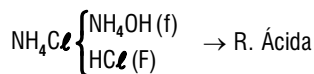
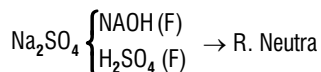
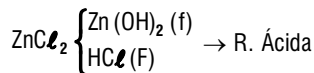
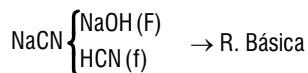
Liberação parcial dos elétrons mais externos e a conseqüente formação de íons positivos.

03. Letra D.

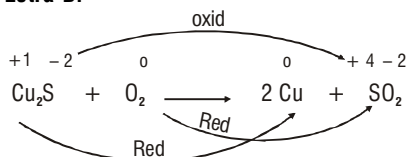
04. Letra C.



05. Letra D.



06. Letra D.



(A) correto: $\text{Cu}^{+1} \rightarrow \text{Cu}^0$

(B) correto: 5

(C) correto: subst. simples

(D) errado: $\text{Cu} + \text{SO}_2$

(E) correto: Cu_2S – sulfeto de cobre I

07. Letra D.

Sal de cozinha: NaCl
composto
iônico

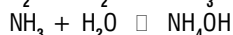
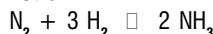
Composto semelhante: KI = iodeto de potássio

08. Letra C.

São sais solúveis em água.

Nitratos e fluoretos de metais alcalinos.

09. Letra A.

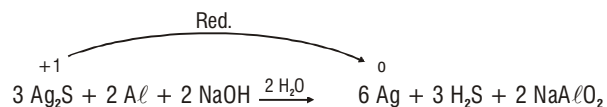


A única fórmula que não se encaixa é N_2O .

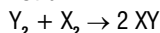
10. Letra C.

Sulfeto de prata: Ag_2S

Hidróxido de sódio: NaOH



11. Letra A.



$$\Delta H = H_p - H_r$$

$$\Delta H = -A - 0$$

$$\Delta H = -A$$

12. Letra B.

I → endotérmicos

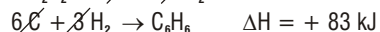
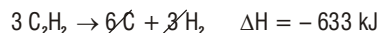
II → exotérmicos

III → exotérmicos

IV → endotérmicos

V → endotérmicos

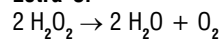
13. Letra D.



14. Letra D.

$$V_m = \Delta n / \Delta t \therefore \frac{18 - 22}{10 - 7} \therefore -4/3 = 1,33 \text{ mol/min.}$$

15. **Letra C.**



$$\begin{cases} 2 \text{ mols H}_2\text{O}_2 & \text{---} & 32 \text{ g O}_2 \\ x & \text{---} & 3,2 \text{ g O}_2 \end{cases}$$

$$x = 0,2 \text{ mol/min.}$$

16. **Letra B.**

E_{at} = energia de ativação para atingir o complexo ativado.

17. **Letra A.**

Reação elementar: $V = k [\text{A}] [\text{B}]^2$

18. **Letra C.**

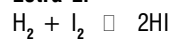
Velocidade independente de B; logo: $[\text{B}]^0$

$$[\text{A}] \times 2 \rightarrow V = 4$$

$$V = [2 \text{ A}]^2$$

$$\text{Logo: } V = k [\text{A}]^2$$

19. **Letra E.**



$$K_c = \frac{[\text{HI}]^2}{[\text{H}_2]}$$

I_2 é sólido, não incorporado ao valor de K_c .

20. **Letra C.**



$$i - 1 \qquad 0 \qquad 0$$

$$\text{eq.: } 1 - 0,47 \qquad 0,47 \qquad 0,47$$

$$K_c = \frac{[\text{PCl}_3][\text{Cl}_2]}{[\text{PCl}_5]} \therefore \frac{0,47 \cdot 0,47}{1 - 0,47} = 0,42 \text{ mol/L}$$